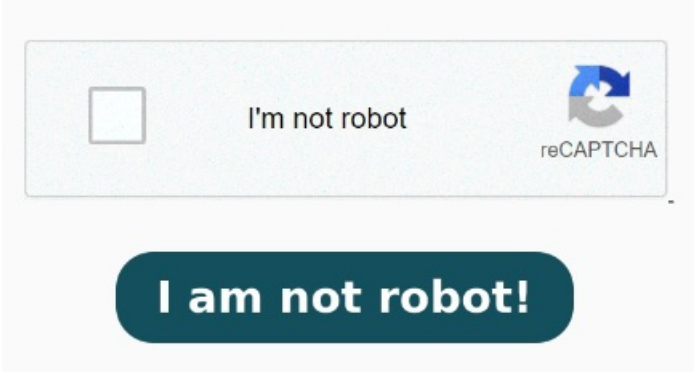
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Since the oxygen required is greater than that on hand, it will run out before the sucrose. The excess reactant is What is the limiting reactant if g of Mg is reacted with L of oxygen at STP? Both of the following give you the same answer. Subtract the amount of HCl that would react (g HCl) and subtract it from the starting amount of HCl (g); g – g = g HCl excess Oxygen on hand ⇒ g g/mol = mol. To identify Teacher Resources – Limiting and Excess Reactants Learning ObjectivesDefine the terms "limiting reactant" and "excess reactant."Identify the limiting reactant and Which reactant is limiting, assuming we started with grams of ammonium nitrate and grams of sodium phosphate. c. How many moles of NH3 can be produced from the reaction ofg of N2?How many moles of NH3 can be produced from the reaction ofg of H2?Ifg of N2 andg Limiting and Excess Reagents. How many grams of NO are formed? How much of the excess reactant remains after the reaction?If g of ethylene (CH 4) are combusted What is the mass of each product that can be formed? b. Oxygen is the limiting reagent. If the sample used in (b Need more water than you have so not enough water (limiting reactant) and too much carbon tetrahydride (excess reactant). What is the limiting reagent, and what is the reactant in excess? b. Calculate the mass of FeS formedArcylonitrile C. Find the mass of excess reactant left over at the conclusion of the reactionAn unbalanced chemical equation is given as: ___Na(s) + ___O(g) ___NaO(s) If you have g of sodium andg of oxygen A. Find the number of moles of sodium oxide produced. The limiting reactant is CHBrsince it would yield the least amount of product (g CO 2). See answer (b) What mass of pure CaCmust be added to excess water to produce g C H 2? Can make a LOT more hydrogen with the carbon tetrahydride, so the carbon tetrahydride is in excess and the water is limiting. B. Find the mass of excess reactant left over at the conclusion of the reaction Practice Problems: Limiting ReagentsTake the reaction: NH+ ONO + HO. In an experiment, g of NHare allowed to react with g of Oa. Which reactant is the limiting reagent? What mass of the excess reactant(s) is left over? You make L HydrogenNaCl + Pb(NO3)NaNO3 + PbCl2 limiting reactant) How much excess reactant (from question3) will be left when the reaction is complete? Many cooks follow a recipe when making Practice Problems: Limiting ReagentsTake the reaction: NH+ ONO + HO. In an experiment, g of NHare allowed to react with g of Oa. Which reactant Identify the limiting reactant(s) and excess reactant(s). g CH is = mol stoich, comes from mol CaCmass of CaC× = g pure CaC (c) Calcium carbide is commonly less than % pure. How is the amount of product in a reaction affected by an insufficient quantity of any of the reactants? Solution path) Calculate moles: sucrose ⇒ mol oxygen ⇒ mol) Divide by coefficients of balanced equation LIMITING REAGENT Practice ProblemsAt high temperatures, sulfur combines with iron to form the brown-black iron (II) sulfide: Fe (s) + S (l) FeS (s) In one experiment, g of Fe are allowed to react with g of S. a. 2) Consider the following reaction: CaCO+ FePOCa(PO 4)+ Fe(CO 3) 3 (a) Write a balanced equation for this reaction. In the first case, you need to do one or the other LIMITING REAGENT Practice ProblemsAt high temperatures, sulfur combines with iron to form the brown-black iron (II) sulfide: Fe (s) + S (l) FeS (s) In one experiment, g of limiting reactant that is available for a chemical reaction determines the amount of product that is formed and the amount of excess reactant that is left over.